

Comments from the reviewer:

Page 13, Line 19. The authors should not generalize. The measurement described in the study by Smith et al. 2016 refers to one Criegee Intermediate, (CH₃)₂COO and, similarly as for the reaction with water, the different Criegee intermediates are likely to decompose at very different rates. For example, CH₂OO is expected to decompose at a rate of ~ 0.6 s⁻¹ (Berndt et al. 2015).

We have revised this statement to clarify that the decomposition rates of different Criegee intermediates are likely different, as suggested.

Page 15, lines 4 to 13. Most probably the main reason is the fact that alpha-pinene has a very low yield of SCI (~0.15) and therefore, for this BVOC, the assumption is correct. One should be careful though to generalize this to all the BVOCs.

We have clarified that a reason for the fact that the majority of OH produced in the a-pinene ozonolysis experiments is due to the rapid decomposition of excited Criegee radicals as the yield of SCIs is relatively low. We have also clarified that this result may not be true for all of the BVOCs tested in these experiments.

Page 15, lines 12 and 13. This sentence is misleading. This reviewer perceives it as if there is uncertainty regarding the mechanisms of production of OH from SCI in atmospheric conditions. This is not correct. It is true that it has not been yet clarified how much of the OH formed during the ozonolysis is generated from excited and how much from stabilized CI, but this will be strongly depending on the BVOC interacting with ozone and larger and more unsaturated VOCs, like ocimene, are expected to produce a larger fraction of SCI from which a large fraction of OH can be expected.

We agree that this sentence is misleading and have removed it to avoid confusion.

Page 15, lines 27-29. It could be worth here adding that the small yield of SCI for alpha-pinene can also be the reason for the good agreement observed despite the simplicity of the model.

As suggested, we have added a statement indicating that because the yield of SCIs from the ozonolysis of alpha-pinene is relatively small, the shortcomings in the treatment of SCIs in the model may not significantly impact the model results.

Page 16, lines 12-14. The majority of the SCI formed during ozonolysis of isoprene consists in CH₂OO (60%, Nguyen et al 2016) that is not forming OH during decomposition. This could explain why no interference was observed.

As suggested, we have added a statement noting the results of the Nguyen et al. study regarding the fact that the majority of SCIs produced in the isoprene ozonolysis mechanism are CH₂OO, which has a low rate of unimolecular reactivity due to its non-syn structure and does not lead to OH formation.

References:

Berndt, T., Kaethner, R., Voigtlander, J., Stratmann, F., Pfeifle, M., Reichle, P., Sipila, M., Kulmala, M., and Olzmann, M.: Kinetics of the unimolecular reaction of CH₂OO and the bimolecular reactions with the water monomer, acetaldehyde and acetone under atmospheric conditions, Phys Chem Chem Phys, 17, 19862-19873, 10.1039/C5CP02224J, 2015.

Smith, M. C., Chao, W., Takahashi, K., Boering, K. A., and Lin, J. J.-M.: Unimolecular Decomposition Rate of the Criegee Intermediate $(\text{CH}_3)_2\text{COO}$ Measured Directly with UV Absorption Spectroscopy, *The Journal of Physical Chemistry A*, 10.1021/acs.jpca.5b12124, 2016.

Nguyen, T. B., Tyndall, G. S., Crouse, J. D., Teng, A. P., Bates, K. H., Schwantes, R. H., Coggon, M. M., Zhang, L., Feiner, P., Miller, D. O., Skog, K. M., Rivera-Rios, J. C., Dorris, M., Olson, K. F., Koss, A., Wild, R. J., Brown, S. S., Goldstein, A. H., de Gouw, J. A., Brune, W. H., Keutsch, F. N., Seinfeld, J. H., and Wennberg, P. O.: Atmospheric fates of Criegee intermediates in the ozonolysis of isoprene, *Phys Chem Chem Phys*, 18, 10241-10254, 10.1039/C6CP00053C, 2016

1 **Measurements of a potential interference with laser-induced** 2 **fluorescence measurements of ambient OH from the ozonolysis** 3 **of biogenic alkenes**

4 Pamela Rickly¹ and Philip S. Stevens^{1,2}

5 ¹School of Public and Environmental Affairs, Indiana, University, Bloomington, IN USA

6 ²Department of Chemistry, Indiana University, Bloomington, IN USA

7 *Correspondence to:* Philip S. Stevens (pstevens@indiana.edu)

8 **Abstract.** Reactions of the hydroxyl radical (OH) play a central role in the chemistry of the atmosphere, and
9 measurements of its concentration can provide a rigorous test of our understanding of atmospheric oxidation.
10 Several recent studies have shown large discrepancies between measured and modeled OH concentrations in
11 forested areas impacted by emissions of biogenic volatile organic compounds (BVOCs), where modeled
12 concentrations were significantly lower than measurements. A potential reason for some of these discrepancies
13 involves interferences associated with the measurement of OH using the Laser-Induced Fluorescence -
14 Fluorescence Assay with Gas Expansion (LIF-FAGE) technique in these environments. In this study, a turbulent
15 flow reactor operating at atmospheric pressure was coupled to a LIF-FAGE cell and the OH signal produced from
16 the ozonolysis of α -pinene, β -pinene, ocimene, isoprene, and 2-methyl-3-buten-2-ol (MBO) was measured. To
17 distinguish between OH produced from the ozonolysis reactions and any OH artefact produced inside the LIF-
18 FAGE cell, an external chemical scrubbing technique was used, allowing for the direct measurement of any
19 interference. An interference under high ozone (between $2-10 \times 10^{13} \text{ cm}^{-3}$) and BVOC concentrations (between
20 approximately $0.1 - 40 \times 10^{12} \text{ cm}^{-3}$) was observed that was not laser generated and was independent of the
21 ozonolysis reaction time. For the ozonolysis of α - and β -pinene, the observed interference accounted for
22 approximately 40% of the total OH signal, while for the ozonolysis of ocimene the observed interference
23 accounted for approximately 70% of the total OH signal. Addition of acetic acid to the reactor eliminated the
24 interference, suggesting that the source of the interference in these experiments involved the decomposition of
25 stabilized Criegee intermediates inside the FAGE detection cell. Extrapolation of these measurements to ambient
26 concentrations suggests that these interferences should be below the detection limit of the instrument.

1 **1. Introduction**

2 The hydroxyl radical (OH) plays an important role in the chemistry of the atmosphere. OH initiates the oxidation
3 of volatile organic compounds (VOCs) which in the presence of nitrogen oxides (NO_x) can lead to the production
4 of ozone and secondary organic aerosols, the primary components of photochemical smog. Because of its high
5 reactivity, measurements of OH can provide a rigorous test of our understanding of the fast radical chemistry in
6 the atmosphere. However, several field campaigns have identified significant discrepancies between measured
7 and modeled OH concentrations, especially in low NO_x forested environments (Rohrer et al., 2014). For example,
8 Ren et al. (2008) found that OH concentrations were well predicted by models to within their combined estimated
9 uncertainty when mixing ratios of isoprene were less than approximately 500 pptv, but measurements acquired in
10 areas with higher mixing ratios of isoprene showed observed OH concentrations that were 3-5 times larger than
11 model predictions. Similarly, measurements in a northern Michigan forest found daytime OH concentrations
12 approximately three times larger and nighttime concentrations 3-10 times larger than model predictions (Tan et
13 al., 2001; Faloon et al., 2001). Aircraft measurements over the Amazon rainforest found OH concentrations to
14 be 40-80% larger than model predictions (Lelieveld et al., 2008), while measurements of OH in the Borneo
15 rainforest were 5-10 times greater than model predictions (Whalley et al., 2011). Similarly, measurements of OH
16 concentrations under high mixing ratios of isoprene in the Pearl River Delta, China were 3-5 times larger than
17 model predictions (Hofzumahaus et al., 2009).

18 Most of these measurements were done using the Laser-Induced Fluorescence – Fluorescence Assay by
19 Gas Expansion (LIF-FAGE) technique. In this technique, ambient air is sampled through an inlet at low pressure,
20 enhancing the OH fluorescence lifetime and allowing temporal filtering of the OH fluorescence from laser scatter
21 (Heard and Pilling, 2003). Fluorescence from OH radicals is distinguished from background scatter and broadband
22 fluorescence through spectral modulation of the laser wavelength. Previous laboratory tests using the Penn State
23 instrument demonstrated that the technique was free from interferences from several species, including spectral
24 interferences from naphthalene, sulfur dioxide, and formaldehyde as well as chemical interferences from high
25 concentrations of H₂O₂, HONO, SO₂, HNO₃, several alcohols and alkanes, propene, and isoprene (Ren et al.,
26 2004). Mixtures of ozone with ethene, propylene and isoprene did not result in any significant signal, suggesting
27 that the ozonolysis of these compounds did not produce an interference in the Penn State instrument, although
28 small interferences were observed with addition of high amounts of ozone and acetone that would be insignificant
29 under ambient conditions (Ren et al., 2004). Measurements of OH in the SAPHIR chamber at atmospheric
30 pressure and under varying concentrations of H₂O, O₃, CO, HCHO, NO, and NO₂ by both an LIF-FAGE
31 instrument and a differential optical absorption spectroscopy (DOAS) instrument were in excellent agreement,

1 suggesting that measurements of OH using the LIF-FAGE instrument were free from artifacts under these
2 conditions (Schlossler et al., 2007). A subsequent intercomparison inside the SAPHIR chamber found that
3 measurements from several different LIF-FAGE instruments under a variety of conditions agreed with each other
4 to within the calibration accuracies of the instruments (Schlosser et al., 2009). No interferences were detected
5 under varying concentrations of O₃, H₂O, NO_x and peroxy radicals, and measurements of OH during the
6 ozonolysis of various mixing ratios of pent-1-ene (6-25 ppb) and trans-2-butene (200 ppb) in approximately 100
7 ppb of ozone also did not reveal a significant interference (Schlosser et al., 2009). In contrast, Hard et al. (2002)
8 observed an interference in their LIF-FAGE instrument during calibrations using the ozonolysis of trans-2-butene
9 under high mixing ratios of both ozone and trans-2-butene. Tests suggested that the interference was not laser
10 generated, and disappeared in air containing 1% water vapor. They suggested that the interference may be due to
11 the dissociation of an intermediate in the ozonolysis mechanism that produces OH in the low-pressure cell of their
12 FAGE instrument (Hard et al., 2002).

13 Recently, Mao et al. (2012) discovered a significant interference associated with their LIF-FAGE
14 measurements of OH in a northern California forest. Using a chemical scavenger to remove ambient OH before
15 air enters the inlet, they found that subsequent spectral modulation revealed a significant amount of internally
16 generated OH from an unknown interference. Measurements using only spectral modulation of the laser
17 wavelength were greater than the measurements when the interference measured using the chemical scavenger
18 was subtracted, and the latter measurements were in better agreement with model predictions (Mao et al., 2012).
19 Similar results were observed by Novelli et al. (2014a) who measured ambient OH concentrations in several forest
20 environments using an external chemical scavenger technique. They found that OH generated inside their
21 detection cell comprised 30-80% of the daytime signal observed using spectral modulation and 60-100% of the
22 signal observed at night. Subtracting the interference from the ambient signal resulted in OH concentrations that
23 were in good agreement with model simulations as well as measurements by a Chemical Ionization Mass
24 Spectrometry (CIMS) instrument (Novelli et al., 2014a).

25 Mao et al. (2012) found that the interference increased with both temperature and measured OH
26 reactivity, suggesting that the interference was related to biogenic emissions in this environment, perhaps the
27 result of BVOC oxidation products entering the sampling cell of the LIF-FAGE instrument which subsequently
28 undergo further reactions and/or decomposition producing the additional OH signal. Novelli et al. (2017) also
29 found that the interference appeared to correlate with temperature similar to the temperature dependence of terpene
30 emissions and was also correlated with the measured ozone concentrations. They concluded that one possible
31 contributor to their interference was the decomposition of stabilized Criegee intermediates inside their instrument.

1 In addition to decomposition leading to OH formation at ambient pressure (Fang et al., 2016; Kidwell et al., 2016;
2 Smith et al., 2017), previous laboratory experiments have shown that BVOC ozonolysis intermediates, such as
3 Criegee intermediates and vinyl hydroperoxides, are likely to promptly decompose to produce OH at low pressures
4 (Kroll et al., 2001a,b). As a result of the large pressure and temperature gradients which occur as the sample enters
5 the FAGE detection cell, the dissociation pathway of these intermediate species may be favored, leading to
6 additional OH production. Recently, Fuchs et al. (2016) performed laboratory and chamber experiments to
7 determine whether the ozonolysis of alkenes produced an OH artifact in their LIF-FAGE instrument. They found
8 that under reactant concentrations that were orders of magnitude greater than ambient, the ozonolysis of propene,
9 α -pinene, limonene, and isoprene produced a detectable interference in their instrument that increased with the
10 turnover rate of the reaction. Extrapolating their results to ambient concentrations of ozone and alkenes would
11 suggest that the ozonolysis of these compounds would not produce a detectable interference in their instrument
12 under ambient conditions (Fuchs et al., 2016).

13 The goal of this work is to determine whether intermediates or products in the ozonolysis of various
14 biogenic alkenes can lead to an interference with OH measurements using the Indiana University LIF-FAGE
15 instrument (IU-FAGE). These experiments focus on the ozonolysis of several biogenic alkenes, including α -
16 pinene, β -pinene, ocimene, isoprene, and 2-methyl-3-buten-2-ol (MBO), all of which have been observed to
17 contribute appreciably to BVOC emissions in forested environments (Guenther et al., 1994; Harley et al., 1998).
18 Measurements of the interference as a function of various instrumental parameters are also provided in an attempt
19 to identify possible sources of the interference and ways it could be minimized.

20 **2. Experimental Section**

21 The ozonolysis experiments were performed using an atmospheric pressure turbulent flow tube with a movable
22 injector similar to that used for ambient measurements of total OH reactivity (Hansen et al., 2014). The 1 m long
23 and 5 cm diameter flow tube was positioned perpendicular to the IU-FAGE detection cell so that the flow would
24 not interfere with the external OH scavenging measurement (section 2.2) (Fig. 1). A Teflon adaptor attached at
25 one end of the flow tube supported the injector, a 1 m stainless steel tube with a 1.25 cm diameter. This injector
26 allowed for the introduction of ozone produced from an ozone generator (Enaly) to the system and could be moved
27 throughout the flow tube to permit varying reaction times between approximately 100 and 420 ms calculated
28 based on the measured velocity in the flow tube. Attached to the end of the injector was a turbulizer used to
29 increase mixing of the reagents at the start of the reaction. A flow of nitrogen of 180 SLPM created a turbulent

1 flow with a Reynolds number of approximately 3750. Ozone concentrations were varied between approximately
2 1 and 3 ppm (2×10^{13} - 7×10^{13} molecules cm^{-3}), with most measurements occurring at ozone concentrations of 2.5,
3 5, and 7×10^{13} molecules cm^{-3} as the ozone generator provided the highest stability at these concentrations. Ozone
4 concentrations in the flow tube were measured using a Teledyne Photometric Ozone Analyzer (model 400E), with
5 an estimated uncertainty of ± 0.5 ppb.

6 BVOC concentrations were introduced into the flow tube by bubbling N_2 through the liquid compound
7 sending the vapor into the reactor. The concentration of the BVOC was estimated from its equilibrium vapor
8 pressure and accounting for dilution into the main flow. Several alkene concentrations were used for each
9 experiment, with approximate concentrations of 2×10^{11} to 4×10^{13} molecules cm^{-3} for α -pinene (Aldrich, 98%),
10 1×10^{11} to 4×10^{13} molecules cm^{-3} for β -pinene (Aldrich, 99%), 9×10^{10} to 5×10^{13} molecules cm^{-3} for ocimene
11 (Aldrich, mixture of isomers, $\geq 90\%$), and approximately 8×10^{13} molecules cm^{-3} for isoprene (Aldrich, 99%), and
12 MBO (Aldrich, 98%). Unfortunately, no direct method for measuring the concentration of these BVOCs was
13 available, and as a result the absolute concentration of BVOCs in the flow tube is highly uncertain due to potential
14 wall losses prior to entering the flow tube.

15

16 **2.1 Detection of OH Radicals**

17 OH radicals were measured using the IU-FAGE instrument, in which ambient air is pulled through either a 0.6 or
18 1 mm diameter nozzle and expanded to a total pressure of approximately 4-9 Torr resulting in a total flow rate of
19 approximately 8-10 SLPM (Dusanter et al., 2008; 2009; Griffith et al., 2013; 2016). Previous field measurements
20 using the IU-FAGE instrument have incorporated a cylindrical inlet (5 cm diameter, 14 cm long) attached to the
21 main detection block resulting in a total distance of approximately 20 cm from the nozzle to the detection volume
22 (Fig. 2). These previous field measurements have utilized both the 0.6 mm nozzle (Griffith et al., 2016) and the 1
23 mm nozzle (Dusanter et al., 2009; Griffith et al., 2013), and the experiments presented here have attempted to
24 reproduce these instrument configurations.

25 The original IU-FAGE laser system used in this study consisted of a Spectra Physics Navigator II
26 YHP40-532Q diode-pumped Nd:YAG laser that produces approximately 5.5W of radiation at 532 nm at a
27 repetition rate of 5 kHz. This laser pumped a Lambda Physik Scanmate 1 dye laser (Rhodamine 640, 0.25 g L^{-1}
28 in isopropanol) that produced tunable radiation around 616 nm, which was frequency doubled to generate 2 to 20
29 mW of radiation at 308 nm. This laser system was recently replaced with a Spectra Physics Navigator II YHP40-
30 532Q that produces approximately 8 W of radiation at 532 nm at a repetition rate of 10 kHz that pumps a Sirah

1 Credo Dye laser (255 mg/L of Rhodamine 610 and 80 mg/L of Rhodamine 101 in ethanol), resulting in 40 to 100
2 mW of radiation at 308 nm. Initial measurements of the OH interference from ocimene and α -pinene ozonolysis
3 were made using the original laser system, with the majority of experiments done using the newer laser system.

4 After exiting the dye laser, the beam is focused onto a 12 m optical fiber to transmit the radiation to the
5 sampling cell where it crosses the expanded air perpendicular to the flow approximately 24 times in a multipass
6 White cell configuration (Fig. 2). The OH molecule is excited and detected using the $A^2\Sigma^+ (v'=0) \leftarrow X^2\Pi_i (v'=0)$
7 transition near 308 nm. A reference cell where OH is produced by thermal dissociation of water vapor is used
8 to ensure that the laser is tuned on-line and off-line of the OH transition to measure the net fluorescence signal.
9 The OH fluorescence is detected by a gated microchannel plate detector (Hamamatsu R5916U-52) and the
10 resulting signal is sent through a preamplifier (Stanford Research SR445) and a photon counter (Stanford Research
11 SRS 400). The detector is switched off during the laser pulse through the use of electronic gating, allowing the
12 OH fluorescence to be temporally filtered from laser scatter. Each offline measurement is recorded for
13 approximately 10 seconds and is averaged and subtracted from the online measurement, averaged for
14 approximately 20 seconds. These measurements are recorded for at least 5 cycles per ozone concentration once
15 the OH concentration has stabilized after several minutes.

16 The sensitivity of the IU LIF-FAGE instrument was calibrated using the UV-water photolysis technique
17 where water vapor is photolyzed to produce known amounts of OH (Dusanter et al., 2008). In these experiments,
18 calibrations were performed under the conditions of the experiments using N_2 , resulting in larger calibration
19 factors compared to ambient air due to fluorescence quenching by oxygen. Nitrogen from liquid boil-off was used
20 instead of synthetic air to reduce the concentration of reactive impurities in the flow tube. The N_2 calibration
21 factors were determined for various instrumental parameters including three cell pressures (4, 7, and 9 Torr), the
22 two nozzle diameters (0.6 mm and 1 mm), and three inlet lengths (3.2 cm, 14 cm, and 24.8 cm) (Table S1, Fig.
23 S1). For reference, previous ambient measurements using the IU LIF-FAGE instrument were typically acquired
24 using the medium inlet length (Fig. 2) and both the 0.6 and 1 mm nozzle diameters, resulting in cell pressures of
25 4-7 Torr depending on the nozzle diameter size and pumping efficiency (Lew et al., 2017a). The error associated
26 with the UV-water photolysis calibration technique is estimated to be $\pm 36\%$ (2σ) (Dusanter et al., 2008)

27 Both nozzle diameters showed similar sensitivities that decreased as the pressure increased due to
28 increased collisional quenching of the OH fluorescence. The calibration factor was also sensitive to the length of
29 the inlet where the ambient air enters the cell and where the OH fluorescence occurs in the detection axis, with
30 the sensitivity decreasing with the increasing inlet length likely due to increased loss of OH radicals on the interior
31 walls of the inlet. For these experiments, the limit of detection was between approximately 5×10^5 - 4×10^6

1 molecules cm^{-3} ($S/N=1$, 10 min integration) depending on the inlet configuration, flow rate, and pressure inside
2 the FAGE detection cell, with the lowest value corresponding to the shortest inlet and lowest pressure, and the
3 highest value corresponding to the longest inlet and highest pressure.

4 **2.2 Measuring the OH Interference**

5 The OH interference was measured using a chemical titration scheme in which perfluoropropylene (C_3F_6 , 99.5%,
6 Matheson) was added through a circular ring surrounding the detection inlet to chemically remove external OH
7 (Griffith et al., 2016) (Fig. 2). Measurements of OH concentrations using spectral modulation compared to
8 chemical modulation with C_3F_6 added externally can reveal OH radicals that are generated inside the detection
9 cell. C_3F_6 was used because it reacts quickly with OH while also having a negligible optical absorption around
10 308 nm (Mao et al. 2012). The injector is approximately 3.5 cm in diameter and 1.0 cm in height and consists of
11 10 equally spaced 0.4 mm holes in the ring that surround the inlet nozzle at the center of the injector. The holes
12 are approximately 0.5 cm in above the level of the detection cell inlet. Increasing the height of the injector above
13 the inlet from 0.5 cm to 4.5 does not impact the overall sensitivity of the instrument, suggesting that OH radicals
14 are not lost on the outer walls of the injector.

15 To determine the flow of C_3F_6 to be used, OH was produced from the photolysis of ambient air using a
16 mercury penlamp placed in front of the inlet (Fig. S2). Once a constant OH signal was established, C_3F_6 was added
17 at varying flows to determine the flow that depleted $\geq 90\%$ of the external OH signal (a total flow of 99.5% C_3F_6
18 of approximately 3-5 sccm). With the flow entering the detection cell inlet of 3-9 SLPM depending on inlet
19 diameter size, this flow of C_3F_6 results in an approximate concentration of $1\text{-}3 \times 10^{16}$ molecules cm^{-3} and a
20 residence time of approximately 0.05 s. This flow rate was able to titrate the externally generated OH regardless
21 of the flow rate set by each individual inlet diameter. To ensure that this flow of C_3F_6 did not titrate OH radicals
22 produced inside the detection cell, a penlamp was placed inside the detection cell directly behind the inlet to
23 generate OH radicals internally. The same external C_3F_6 flow was again introduced to ensure that the concentration
24 of C_3F_6 after expansion into the detection cell was not high enough to titrate any internally generated OH.

25 By applying this C_3F_6 titration method to the alkene ozonolysis experiments, any OH produced in the
26 flow tube was expected to be removed. Although the ozonolysis reactions may still lead to additional OH
27 production during the residence time in the injector, kinetic simulations suggest that under the conditions of these
28 experiments the steady-state concentration of OH in the presence of C_3F_6 would be near or below the detection
29 limit of the instrument (less than 5×10^5 molecules cm^{-3}). However, for some of the high concentration
30 experiments, the amount of C_3F_6 added may not have been sufficient to reduce the steady-state concentration of

1 OH to below the detection limit of the instrument, especially for the ocimene experiments due to the high reactivity
2 of ocimene with ozone. However, the model simulations suggest that even for these high concentration
3 experiments the remaining steady-state OH concentrations represented less than 10% of the observed interference.
4 Thus, the majority of OH that was measured using spectral modulation after C₃F₆ addition would be an
5 interference generated internally. Subtraction of this interference from the measurement acquired before C₃F₆
6 addition should reflect the steady-state OH concentration generated in the flow tube, which can then be compared
7 to literature values of the OH yield for the ozonolysis of these alkenes.

8 3. Results and Discussion

9 3.1 Ozonolysis experiments

10 The interference tests were performed during the alkene ozonolysis experiments with a reaction time in the flow
11 tube of approximately 420 ms, a reaction time longer than that required for the system to reach steady-state (less
12 than approximately 20 ms based on model calculations), and the results are shown in Figs. 3 and 4. In these
13 figures, the open symbols are the measured OH concentration produced from the ozonolysis reaction without
14 addition of C₃F₆ (With interference), and the filled symbols represent the OH signal after the signal measured with
15 C₃F₆ addition is removed (Without interference) using the 0.6 mm nozzle, shown on the same plot to illustrate the
16 magnitude of the interference. The expected steady-state OH concentrations based on previous measurements of
17 the OH yield for each compound are also shown in each figure by the solid line, calculated using the following
18 equation:

$$19 \quad [\text{OH}]_{ss} = \frac{k_{O_3+VOC}\alpha[O_3][alkene]}{k_{OH+VOC}[alkene] + k_{OH+O_3}[O_3] + k_{wall}} \approx \frac{k_{O_3+VOC}\alpha[O_3]}{k_{OH+VOC}} \quad (1)$$

20 In this equation, k_{O_3+VOC} is the rate constant for the O₃ + alkene reaction with an OH yield of α , k_{OH+VOC} is the rate
21 constant for the OH + alkene reaction, k_{OH+O_3} is the rate constant for the OH + O₃ reaction, and k_{wall} is the first-
22 order loss of OH on the walls of the reactor, measured as described in Hansen et al. (2014). Determining the
23 expected steady-state OH concentration from this equation has the advantage of being independent of the
24 concentration of each BVOC. Rate constants for the OH and O₃ reactions were obtained from recommendations
25 by Atkinson (1997) and Atkinson et al. (2006). For these calculations, the loss of OH from reaction with ozone
26 (approximately 7 s⁻¹) as well as wall loss of OH (approximately 3.6 s⁻¹), were neglected, as they were much smaller
27 than loss of OH due to reaction with the alkenes due to the high concentration of alkenes used in these experiments.

1 As illustrated in Fig. 3, the measured OH concentration from the ozonolysis of α -pinene without C_3F_6
2 addition is consistently greater than the measurements after the interference is subtracted, indicating that a
3 significant concentration of OH is being produced inside the detection cell. In these figures, instabilities in the
4 concentration of ozone produced by the ozone generator resulted in small variations in the measured ozone
5 concentrations between measurements with and without added C_3F_6 . Experiments without α -pinene in which
6 ozone alone was sampled by the detection cell showed a negligible OH signal, suggesting that the interference
7 was related to the presence of both the alkene and ozone. The observed interference measured with the addition
8 of C_3F_6 accounted for approximately 40-60% of the observed signal in the absence of C_3F_6 . When the measured
9 interference was subtracted from the overall signal, the resulting OH concentrations were in reasonable agreement
10 with the OH yield expected from α -pinene ozonolysis resulting in an OH yield of approximately 0.81 ± 0.10 ,
11 determined from a weighted fit of the slope of the plot of the OH concentration versus ozone concentration, in
12 good agreement with the value of 0.76 ± 0.11 as reported by Chew and Atkinson (1996) and the value of $0.91 \pm$
13 0.23 reported by Siese et al. (2001).

14 The measured interference appeared to depend on the length of the inlet, with the greatest interference
15 observed with the longest inlet (Fig. 3), while the interference measured with the short and medium length inlets
16 were comparable. Similar results were observed with the use of the 1 mm nozzle diameter (Fig. S3), with the
17 longest inlet displaying the greatest interference. These results are similar to that observed by Fuchs et al. (2016),
18 who found that the interference from the ozonolysis of α -pinene, limonene and isoprene also increased with the
19 length of the inlet in their FAGE instrument. But while the inlet length appears to exhibit a trend in the OH
20 interference, the cell pressure did not appear to significantly impact the level of the interference except at the
21 longest inlet length. These results are in contrast to the results of Fuchs et al. (2016), who found that the
22 interference in their instrument decreased with cell pressure, although the effect was greater with the longer inlet.
23 The different cell pressures in these experiments were obtained by changing the pumping speed, which impacts
24 the velocity and residence time of the airstream inside the detection cell. This suggests that increasing the reaction
25 time for the short and medium length inlets does not significantly impact the interference, while increasing the
26 reaction time for the longest inlet does increase the interference. However, at the longest inlet length, the increased
27 reaction time likely also leads to increased collisions with the interior surfaces of the detection cell, which could
28 also lead to increased dissociation and production of OH. Increasing the nozzle diameter decreases the flow
29 velocity inside the detection cell, which increases the reaction time and the frequency of wall collisions using the
30 longer inlet length. Although this also increases the loss of OH generated by the interference, it appears that the
31 rate of production of OH by the interference is greater than the loss rate of OH on the walls of the inlet. These

1 results appear to be independent of the α -pinene concentration (Fig. 4, top) as the level of the interference is
2 similar for estimated α -pinene concentrations between approximately 1 and 4×10^{13} molecules cm^{-3} .

3 The results from the ozonolysis of β -pinene are also shown in Fig. 4 (middle). Similar to that observed
4 for the ozonolysis of α -pinene, the observed interference measured with the addition of C_3F_6 accounted for
5 approximately 40% of the observed signal in the absence of C_3F_6 . When the measured interference was subtracted
6 from the overall signal, the measured OH concentrations results in an overall OH yield of approximately $0.40 \pm$
7 0.01 , which is in good agreement with the value reported by Atkinson et al. (1992) of 0.35 ± 0.05 . The percentage
8 of interference as well as the resulting OH yield was relatively constant for each β -pinene concentration used;
9 however, the results suggest that the OH yield may increase with increasing concentrations of β -pinene.

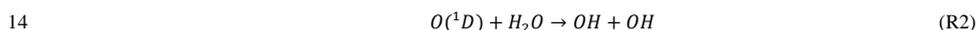
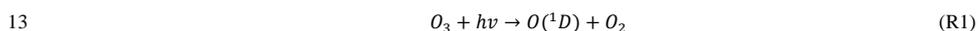
10 Measurements of the concentration of OH produced from the ozonolysis of various concentrations of
11 ocimene are also shown in Fig. 4 (bottom). In contrast to the ozonolysis of α -pinene, the interference and the
12 observed OH yield appear to increase with increasing ocimene concentration. At ocimene concentrations of
13 approximately 4×10^{11} molecules cm^{-3} , an OH yield of approximately 0.28 ± 0.02 was measured. However, the
14 apparent OH yield increased to a yield of 1.0 ± 0.03 at the highest ocimene concentration (approximately 5×10^{13}
15 molecules cm^{-3}). Additional ocimene ozonolysis experiments were performed using the two nozzle diameters and
16 three inlet lengths with the results shown in Fig. S4. These results show that the measured OH yield from ocimene
17 ozonolysis after subtracting the interference is only consistent with the results of Aschmann et al. (2002) of 0.55
18 ± 0.09 for an ocimene concentration of approximately 1×10^{13} molecules cm^{-3} . The increase in the apparent OH
19 yield and interference as the concentration of ocimene was increased may be due to the additional ozonolysis of
20 the reaction products, as the products of ocimene ozonolysis are unsaturated. The higher ocimene concentrations
21 would lead to increased concentrations of reaction products, which in turn could also contribute to the measured
22 interference and OH yield. Additional experiments are needed in order to better quantify the OH yield from the
23 ozonolysis of ocimene.

24 In contrast to the measurements involving α -pinene, β -pinene, and ocimene, similar ozonolysis
25 experiments involving isoprene and MBO did not produce a detectable OH signal or interference under the highest
26 ozone and alkene concentrations. These results may reflect the lower relative reactivity of isoprene and MBO with
27 O_3 compared to reaction with OH as well as a lower yield of OH under the conditions of these experiments.
28 Simulations of the kinetics using the MCM mechanism for the ozonolysis of isoprene result in steady-state OH
29 concentrations that are a factor of 50 lower than that predicted for the ozonolysis of α -pinene, consistent with a
30 lower steady-state OH concentration estimated using Equation 1, and are near or below the detection limit of the

1 instrument. However, these results do not rule out the possibility of an interference from the ozonolysis of these
2 compounds. The lower reactivity of isoprene and MBO with ozone would likely require longer reaction times or
3 higher concentrations of ozone and/or alkenes to establish steady-state OH concentrations above the detection
4 limit of the instrument in these experiments. Fuchs et al. (2016) did find a significant unexpected OH signal in
5 laboratory under very high concentrations of isoprene and a reaction time on the order of 1 s. Future experiments
6 will involve higher concentrations of ozone, isoprene and MBO and/or longer reaction times to increase the
7 ozonolysis rate for these reactions and to determine the potential magnitude of the interference from isoprene and
8 MBO in the IU-FAGE instrument.

9 **3.2 Interference measurements as a function of laser power, ozone, and ozonolysis reaction time**

10 A potential interference with the detection of OH radicals by laser-induced fluorescence is the production of OH
11 through photolysis of ozone by the laser, with the resulting excited oxygen atoms reacting with ambient water
12 vapor to produce OH (reactions R1 and R2) (Stevens et al., 1994; Ren et al., 2004):



15 The resulting OH signal from this two-photon process would display a quadratic dependence on laser power. One
16 possible source of the interference observed in the ozonolysis experiments could be O(¹D) produced from the
17 photolysis of ozone by the laser reacting with the alkene leading to the production of OH radicals through a
18 hydrogen abstraction mechanism. The OH signal from this two-photon process would also display a quadratic
19 dependence on laser power.

20 To confirm that laser generated OH was not occurring within this instrument and was not the source of
21 the interference, experiments were conducted over a wide range of laser powers (0.6-12 mW) to determine
22 whether the observed interference displayed a dependence on laser power consistent with a laser-generated
23 mechanism. Fig. 5 (top) shows the measurements of the OH concentration with and without the interference for
24 the ozonolysis of ocimene as a function of laser power. These experiments were done under dry conditions to
25 minimize potential laser-generated interferences from reactions R1 and R2. Because the OH concentration
26 remained relatively stable over the range of laser powers used in these experiments, there is no indication that the
27 interference was due to a laser-generated mechanism. These results are similar to that found by Novelli et al.
28 (2014a) and Fuchs et al. (2016).

29 A plot of the interference at different ozone concentrations is also shown in Fig. 5 (middle) for the
30 ozonolysis of α -pinene as an example. Comparing the measured interference for each inlet length and ozone

1 concentration shows a trend in which the interference appears to increase with increasing ozone concentration,
2 with the greatest increase occurring with the longest inlet. Similar to the results of Fuchs et al. (2016), the observed
3 interference appears to increase with the overall ozonolysis turnover time, defined as the steady-state rate of OH
4 radical propagation, expressed here as the rate of alkene ozonolysis (Fig. 6). Although the magnitude of the
5 interference observed here is greater due to the greater turnover rates used in these experiments, the slope of the
6 observed interference using the short inlet as a function of the turnover rate appears to be similar to that observed
7 by Fuchs et al. (2016), although the results presented here are highly scattered, likely due to the large uncertainty
8 associated with estimating the concentration of each BVOC. However, the level of the observed interference is
9 greater than that illustrated in Fig. 6 for the measurements using the long inlet, suggesting the similarity with the
10 results of Fuchs et al. (2016) may be fortuitous, as differences in the design of the instrument impacts the level of
11 the interference. In contrast to the results of Fuchs et al. (2016), the level of interference as a function of the
12 turnover rate is not similar for all of the BVOCs tested. While the observed interference as a function of turnover
13 rate appears to be similar for the ozonolysis of α -pinene and ocimene, the observed interference for the ozonolysis
14 of β -pinene is significantly less. This may also be related to uncertainties associated with estimates of the
15 concentration of β -pinene in the flow tube, but may also suggest differences in the mechanism for the production
16 of the interference for some BVOCs. Additional measurements of the interference for other BVOCs are needed
17 to resolve this discrepancy.

18 The interference signals from the ozonolysis of α -pinene, β -pinene, and ocimene for each experimental
19 condition are summarized in Table S2. The interference from α -pinene expressed as a percentage of the total OH
20 signal at several ozone concentrations and cell pressures for the short, medium, and long inlets are shown in Fig.
21 S5. On average, the percent interference was similar for ozone mixing ratios between 1-3 ppm, with values of
22 approximately $45 \pm 3\%$, $37 \pm 8\%$, and $58 \pm 8\%$ for the short, medium, and long inlets, respectively for the 1 mm
23 nozzle, and approximately $54 \pm 7\%$, $50 \pm 13\%$, and $65 \pm 11\%$ for the short, medium, and long inlets, respectively
24 for the 0.6 mm nozzle. As discussed above, the similarity of the measured interference between the short and
25 medium length inlets suggest that the level of interference is not directly related to the residence time inside the
26 FAGE detection cell, but may be the result of increased collisions with the interior surfaces of the detection cell
27 that occurs when using the longest inlet. Similar results were observed for the ozonolysis of ocimene (Fig. S6),
28 although additional experiments will be needed to determine whether a similar trend with inlet length and
29 residence time is observed.

30 The interference was also measured over varying reaction times within the flow tube, and an example of
31 the results is shown in Fig. 5 (bottom). The results appear to indicate that the level of interference does not depend

1 on the ozonolysis reaction time in the flow tube. If the interference was due to a stable oxidation product, it would
2 be expected to increase with reaction time as the concentration stable oxidation products accumulated in the flow
3 tube. These results suggest that the interference is not due to a stable oxidation product but may instead be due to
4 a steady-state intermediate in the ozonolysis mechanism. The consistency of the measured interference relative to
5 the OH concentration produced from the ozonolysis mechanism may suggest that the source of the interference is
6 related to the source of OH, such as the Criegee intermediate.

7 3.3 Stabilized Criegee intermediates as a source of the interference

8 **3.3.1 Decomposition of Criegee intermediates produced inside the FAGE detection cell as a source of**
9 **the interference.** The ozonolysis of alkenes involves the addition of ozone to the double bond, forming a primary
10 ozonide which quickly decomposes to an excited Criegee intermediate and a carbonyl compound. Depending on
11 its structure, the excited Criegee intermediate can either decompose to form an OH radical, or can be stabilized
12 by collisions forming a stabilized Criegee intermediate (SCI), which can also thermally dissociate into OH at
13 longer reaction times (approximately 1 sec for tetramethylethylene, Kroll et al., 2001b). At shorter reaction times,
14 the excited Criegee intermediate is in steady-state with respect to dissociation and stabilization, resulting in OH
15 concentrations that are generally independent of reaction time, but increase at longer reaction times due to the
16 additional OH production from SCIs (Kroll et al., 2001b). Rate constants for decomposition and stabilization of
17 excited Criegee intermediates are estimated to be on the order of 10^4 - 10^9 s⁻¹ (Horie and Moortgat, 1991), while
18 the rate constant for decomposition of the stabilized (CH₃)₂COO Criegee intermediate, has been measured to be
19 on the order of 10² s⁻¹ (Smith et al., 2016). Although different Criegee intermediates will likely decompose at
20 different rates, one might expect that the concentration of excited Criegee intermediates to be very low compared
21 to the concentration of SCIs, many of which have estimated lifetimes on the order of milliseconds relative to
22 decomposition. The atmospheric fate of SCIs will depend on the rate of decomposition relative to reaction with
23 other atmospheric trace gases such as SO₂ or H₂O among others (Novelli et al., 2014b), which are also likely
24 different for different intermediates. This lifetime may be longer in the low pressure region of the FAGE detection
25 cell due to the decrease in concentration of reactants such as SO₂ or H₂O, allowing these SCIs to collide with the
26 walls of the detection cell.

27 The time dependence of the observed interference in these experiments, with the interference appearing
28 to be independent of reaction time except under the longest reaction time, at first appears to be consistent with a
29 mechanism that involves the formation of excited Criegee radicals inside the FAGE detection cell. However, it is
30 unlikely that excited Criegee intermediates could be produced directly from alkene ozonolysis inside the IU-

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1 FAGE detection cell as the decrease in reactant concentrations from the expansion to low pressure leads to
2 turnover rates that are approximately two orders of magnitude smaller than that shown in Fig. 6. For the ozonolysis
3 of α -pinene, a reaction time of approximately 0.5-1 s would be required to internally produce the observed OH
4 signals in these experiments, which is much longer than the reaction time inside the detection cell (on the order
5 of 1-2 ms). The short reaction time in the detection cell is also too short for OH concentrations to reach steady-
6 state. The lower pressure in the FAGE detection cell may result in higher yields of OH due to the lower rate of
7 stabilization of the excited Criegee intermediate under these conditions (Kroll et al., 2001a). Nevertheless, an OH
8 yield of one would still require a reaction time of approximately 0.18 to 0.35 s to produce the observed internal
9 OH concentrations.

10 **3.3.2 Decomposition of Criegee intermediates produced outside of the FAGE detection cell as a source**
11 **of the interference.** Previous measurements have demonstrated that the stabilized Criegee intermediates produced
12 external to the FAGE detection cell from the ozonolysis of propene and (*E*)-2-butene can decompose to produce
13 OH radicals upon entering the low pressure region of the FAGE detection cell (Novelli et al., 2014b). The time
14 dependence of the OH production was consistent with previous measurements of the rate of unimolecular
15 decomposition of SCIs (Novelli et al., 2014b). To determine whether the decomposition of stabilized Criegee
16 intermediates are the source of the interference in these experiments, acetic acid was added to the flow tube as an
17 SCI scavenger. Welz et al. (2014) have reported direct measurements of the rate constants for the reactions of
18 CH_2OO and CH_3CHOO Criegee intermediates with formic and acetic acids and found values in excess of 1×10^{10}
19 $\text{cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$. In contrast, the rate constant for reaction of acetic acid with OH is approximately 7×10^{13}
20 $\text{cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$ (Atkinson et al., 2006). Acetic acid was chosen over SO_2 or water as an SCI scavenger as SO_2
21 fluorescence at 308 nm can interfere with OH measurements (Fuchs et al., 2016), while the addition of water in
22 the presence of the high concentrations of ozone in these experiments would have led to significant laser-generated
23 OH from reactions R1 and R2. Assuming that other Criegee intermediates react similarly, addition of acetic acid
24 should scavenge SCIs in the flow tube. If SCIs are the source of the OH interference, the measured OH signal
25 resulting from addition of acetic acid should be equivalent to the OH signal when the OH interference measured
26 using external C_3F_6 titration is subtracted. For these experiments, approximately 9×10^{12} molecules cm^3 of acetic
27 acid was introduced into the flow tube, and allowed to react for approximately 200 ms. At this concentration, the
28 reaction with acetic acid was modeled to have a minimal impact on the steady-state concentration of OH.

29 The results of these experiments for the ozonolysis of α -pinene and ocimene are shown in Fig. 7. In this
30 figure, the open symbols are the measured total OH signal due to OH produced in the flow tube from both excited
31 and stabilized Criegee intermediates plus the OH interference, while the solid red symbols represent the remaining

1 OH signal after the interference (measured after external C₃F₆ addition) is subtracted. The solid green symbols
2 represent the measured OH signal (without external C₃F₆ addition) when acetic acid is added to the flow tube. As
3 can be seen from this figure, the measured OH signal when acetic acid is added is similar to the OH signal when
4 the interference is subtracted, suggesting that the interference is due to the decomposition of SCIs produced from
5 the ozonolysis of α -pinene inside the IU-FAGE detection cell, consistent with the results of Novelli et al. (2014b).
6 These results also suggest that the majority of the OH radicals in the flow tube are produced from the rapid
7 decomposition of excited Criegee radicals, consistent with the low yield of SCIs produced from the ozonolysis of
8 α -pinene (Drozd and Donahue, 2011). If stabilized Criegee intermediates were responsible for a significant
9 fraction of the OH produced in the flow tube, it is likely that addition of acetic acid would have led to measured
10 OH signals that were significantly lower than the OH signal after the interference was subtracted. However, these
11 results do not exclude the possibility that SCIs from the ozonolysis of the other BVOCs in these experiments are
12 also thermally decomposing and contributing to OH production in the flow tube. Recent measurements suggest
13 that the unimolecular decomposition of the (CH₃)₂COO Criegee intermediate can be rapid and compete with
14 reaction of this intermediate with water or SO₂ (Smith et al., 2016; Chhantyal-Pun et al., 2017). A similar rate for
15 the decomposition of SCIs in the experiments reported here could also compete with reaction of this intermediate
16 with acetic acid. ▼

17 To determine whether the magnitude of the interference observed in these experiments was consistent
18 with the concentration of stabilized Criegee intermediates in the flow tube, the concentration of these
19 intermediates was estimated from a chemical model of the ozonolysis of α -pinene using the Master Chemical
20 Mechanism (MCM v3.2, Jenkin et al., 1997; Saunders et al., 2003). In this mechanism, the ozonolysis of α -pinene
21 results in the formation of two Criegee intermediates, APINOOA and APINOOB. APINOOA decomposes to OH
22 radicals through two channels with different co-products with a decomposition rate constant of $1 \times 10^6 \text{ s}^{-1}$.
23 APINOOB also decomposes to form an OH radical with a decomposition rate constant of $0.5 \times 10^6 \text{ s}^{-1}$, but can
24 also form a stabilized Criegee intermediate (APINBOO) with a rate constant of $0.5 \times 10^6 \text{ s}^{-1}$. In addition to reaction
25 with CO, NO, NO₂, and SO₂ (which are not significant in these experiments), the stabilized Criegee intermediate
26 in this mechanism can also decompose to form pinonaldehyde or pinonic acid, with rate constants that depend on
27 the concentration of water vapor ($1.40 \times 10^{-17} \text{ s}^{-1} \text{H}_2\text{O}$ and $2.00 \times 10^{-18} \text{ s}^{-1} \text{H}_2\text{O}$, respectively). This version of the
28 MCM does not include a mechanism for the formation of OH from the stabilized Criegee intermediate, nor does
29 it distinguish between the *syn* and *anti* isomers of the Criegee intermediates, which may have different
30 decomposition rates and yields. Despite its potential shortcomings, the mechanism was used for simplicity to
31 provide a rough estimate of the concentration of Criegee intermediates to compare with the experimental

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1 measurements. Given that the yield of SCIs from the ozonolysis of α -pinene is relatively small (Drozd and
2 Donahue, 2011), these shortcomings in the treatment of SCIs in the model may not significantly impact the model
3 results.

4 Under the conditions of these experiments, the model predicts a steady-state concentration of stabilized
5 Criegee intermediates (APINBOO) of approximately $2\text{-}6 \times 10^8$ molecules cm^{-3} (Fig. 7). Based on these results,
6 the observed OH interference in these experiments could be explained if approximately 5% of these intermediates
7 dissociated and produced OH radicals inside the IU-FAGE detection cell, assuming that the transmission of these
8 stabilized Criegee intermediates through the inlet is similar to that for OH. In contrast, the MCM predicted steady-
9 state concentration of excited Criegee radicals (APINOOA and APINOOB) in the flow tube was calculated to be
10 on the order of 10^3 molecules cm^{-3} , much less than the observed interference. Thus it is unlikely that excited
11 Criegee intermediates are the source of the interference, but rather the interference is the result of decomposition
12 of SCIs in the FAGE detection cell. These results are in contrast to the results of Fuchs et al. (2016), who found
13 that in their ozonolysis experiments that the addition of SO_2 as a scavenger for Criegee intermediates did not affect
14 the observed OH signal. They concluded that SCIs were not the cause of the interference in their instrument,
15 although they could not rule out that the interference was due to the decomposition of particular SCI isomers that
16 did not react with SO_2 (Fuchs et al., 2016).

17 If stabilized Criegee intermediates are the source of the interference in these measurements, one might
18 have expected to observe an interference associated with the ozonolysis of isoprene even though the expected
19 yield of OH was below the detection limit of the instrument. Simulations using the Master Chemical Mechanism
20 suggest that the concentration of SCIs in the isoprene ozonolysis mechanism are similar to the concentration of
21 SCIs in the α -pinene ozonolysis mechanism. Given that OH yield from the decomposition of excited CIs in the
22 isoprene mechanism is lower than that for the α -pinene mechanism, the absence of a detectable interference in
23 the isoprene experiments described here may suggest that the decomposition of these intermediates inside the
24 FAGE detection cell may also result in a lower OH yield. These results are also consistent with the recent study
25 by Nguyen et al. (2016), who found that the majority of SCIs produced in the ozonolysis of isoprene were the
26 simplest SCI, CH_2OO . The unimolecular reactivity of CH_2OO is greatly reduced compared to other SCIs due to
27 the fact that it is non-syn, without any facing hydrocarbon groups (Nguyen et al., 2016), and does not lead to OH
28 formation. Consistent with the observation that the observed interference appears to be a constant fraction of the
29 total OH yield in these experiments independent of the ozone concentration and the turnover rate (Figs. S5 and
30 S6), the observed interference for each alkene is likely proportional to the OH yield from the ozonolysis
31 mechanism, with the absence of an observed interference in the isoprene experiments consistent with the lower

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1 OH yield in the isoprene ozonolysis mechanism. Additional experiments measuring the interference from the
2 ozonolysis of isoprene is needed to resolve this issue.

3 **3.3.3 Interference from Criegee intermediate decomposition on LIF-FAGE calibrations.** These results
4 also imply that using the ozonolysis of alkenes as a potential source for calibrating LIF-FAGE instruments may
5 lead to an overestimation of the instrument sensitivity given the potential for this source to produce an interference.
6 As mentioned above, Hard et al. (2002) observed an interference in their LIF-FAGE instrument during calibrations
7 using the ozonolysis of trans-2-butene under high mixing ratios of both ozone (up to 28 ppm) and trans-2-butene
8 (greater than 12 ppb). They found that the interference disappeared in the presence of 1% water vapor. Given the
9 potentially rapid reaction of Criegee intermediates with water vapor (Chao et al., 2015), this may suggest that the
10 source of their interference was also the decomposition of Criegee intermediates inside their detection cell.
11 However, Dusanter et al. (2008) found that the instrument sensitivities derived from the ozone-alkene calibration
12 technique using trans-2-butene were systematically lower than those derived from the water-vapor UV- photolysis
13 technique, in contrast to what might be expected if the measurements from the ozonolysis technique were impacted
14 by an interference from Criegee intermediates. But because these ozonolysis experiments were done under humid
15 conditions, it is possible that the Criegee intermediates were scavenged by water vapor prior to entering the IU-
16 FAGE detection cell, thus minimizing the interference. Although the rate constant of Criegee intermediates with
17 water vapor depends on the structure of the intermediate (Vereecken et al., 2015), future calibrations of LIF-
18 FAGE instruments using the ozonolysis of alkenes should be done in the presence of water vapor or another
19 Criegee scavenger to insure that the calibration is not influenced by this potential interference.

20 **4. Atmospheric Implications**

21 The percent interference observed in these studies (Figs. S5, S6) is similar to the ambient measurements reported
22 by Mao et al. (2012), who measured an interference that was approximately 50% of the total OH signal in an
23 environment dominated by 2-methyl-3-buten-2-ol (MBO), monoterpenes, sesquiterpenes, and related oxygenated
24 compounds. Novelli et al. (2014a) also found that in their measurements of OH in several forest environments the
25 interference accounted for 30-80% of their total OH signal during the daytime. External addition of SO₂ as an SCI
26 scavenger during some of these measurements resulted in the complete removal of the interference, suggesting
27 that the source of the interference was the decomposition of ambient SCIs inside their detection cell (Novelli et
28 al., 2017). However, the magnitude of the interference signal relative to the calibration for OH was much greater
29 than the expected concentration of SCIs in these environments, which was estimated to be on the order of 5×10^4

1 molecules cm^{-3} (Novelli et al., 2017). One possible reason for this discrepancy is a greater sensitivity of the LIF-
2 FAGE instrument to the detection of ambient SCIs relative to ambient OH, perhaps due to a greater transmission
3 efficiency of SCIs into the FAGE detection cell. Tests to measure the sensitivity of their LIF-FAGE instrument to
4 the detection of SCIs relative to OH found that the transmission of *syn*- CH_3CHOO through different nozzle
5 designs was different from the transmission of OH radicals (Novelli et al., 2017). These results suggest that the
6 sensitivity of their LIF-FAGE instrument may be different for the detection of SCIs and that using the calibration
7 factor for OH radicals to estimate the SCI concentration from the interference may not be appropriate. However,
8 the sensitivity of their LIF-FAGE instrument to detection of SCIs would have to be a factor of 100 greater than
9 that for OH based on the estimated concentration of ambient SCIs (Novelli et al., 2017). Previous measurements
10 on a similar LIF-FAGE instrument demonstrated that the measured OH signal is relatively insensitive to the shape
11 and coating of the inlet (Stevens et al., 1994). In addition, external calibrations including inlet losses were found
12 to be similar to internal calibrations that did not use the inlet. These results imply that heterogeneous loss of OH
13 radicals on the inlet is not occurring to a significant extent (Stevens et al., 1994).

14 Similar to that observed by Fuchs et al. (2016), extrapolating the interference observed in the experiments
15 presented here as a function of the ozonolysis turnover time (Fig. 6) to concentrations and turnover rates typically
16 observed in the atmosphere (approximately 1.5 ppb hr^{-1} , Hakola et al., 2012; Fuchs et al., 2016) would suggest
17 that the interference under ambient conditions would be near the detection limit of the IU-FAGE instrument
18 (approximately $4 \times 10^5 \text{ molecules cm}^{-3}$). Consistent with this result, previous measurements of OH radical
19 concentrations by the IU-FAGE instrument in an urban environment during CalNex (California Research at the
20 Nexus of Air Quality and Climate Change) using the external chemical titration technique found no evidence of
21 an unknown interference (Griffith et al., 2016). Measurements of OH by the IU-FAGE instrument in a forested
22 environment in northern Michigan during and after the CABINEX (Community Atmosphere-Biosphere
23 Interactions Experiment) campaign using the external chemical titration technique also did not reveal any
24 unknown interferences (Griffith et al., 2013). However, the measurements during CABINEX were done under
25 relatively cool conditions and low ozone concentrations, where daytime maximum ozone mixing ratios were
26 approximately 30 ppb on average. The relatively cool conditions suggest that mixing ratios of BVOCs may have
27 been relatively low during this campaign, as mixing ratios of isoprene were less than 2 ppb. During CalNex,
28 maximum daytime mixing ratios of ozone were higher (approximately 50-60 ppb on average during the week and
29 approximately 70-80 ppb on average on the weekends), but mixing ratios of isoprene were even lower, less than
30 1 ppb on average, (Griffith et al., 2016), suggesting that the mixing ratio of other BVOCs may also have been
31 lower.

1 In contrast, recent measurements of OH concentrations by the IU-FAGE instrument in an Indiana forest
2 using the external chemical titration technique did reveal an interference that correlated with both temperature
3 and ozone concentrations, similar to the results of Mao et al. (2012) (Lew et al., 2017b). Compared to the
4 conditions during CABINEX, these measurements were done under higher daytime maximum ozone mixing ratios
5 (approximately 40 ppb) and relatively warmer conditions, where isoprene mixing ratios greater than 4 ppb suggest
6 higher mixing ratios of other BVOCs. The magnitude of the observed interference (on the order of 10^6 molecules
7 cm^{-3}) was similar to that observed previously by other LIF-FAGE instruments (Mao et al., 2012; Novelli et al.,
8 2014a; 2017), and accounted for approximately 60% of the total measured OH signal on average. The known
9 interference from laser-generated OH varied with laser power, ambient ozone and water concentrations, but was
10 approximately half of the total measured interference on some days, while on other days accounted for all of the
11 measured interference. However, similar to that observed by Novelli et al. (2017), estimates of the ambient
12 concentration of SCIs on the order of approximately $4\text{-}5 \times 10^4$ molecules cm^{-3} for similar environments (Percival
13 et al., 2013; Novelli et al., 2017) suggest that the observed interference in these measurements may not be solely
14 due to ambient SCIs unless there are other significant sources of Criegee radicals that are not accounted for in
15 these models.

16 These results suggest that although SCIs may be contributing to the observed interference, there may
17 exist an unknown interference in these measurements that also correlates with the concentration of ozone and
18 BVOCs, perhaps due to oxidation products of BVOCs not tested in the experiments reported here (Fuchs et al.,
19 2016; Novelli et al., 2017). Recently, Fuchs et al. (2016) reported an interference in their LIF-FAGE instrument
20 associated with NO_3 radicals, although the exact mechanism of the interference remains unknown. These results
21 suggest that there may be other potential interferences associated with the technique in addition to the
22 decomposition of SCIs that involve complex homogeneous and/or heterogeneous mechanisms inside the FAGE
23 detection cell.

24 **5. Summary and Conclusions**

25 Measurements of OH concentrations produced from the ozonolysis of α -pinene, β -pinene, and ocimene revealed
26 a potential interference associated with the Indiana University LIF-FAGE instrument. The observed interference
27 did not appear to be laser generated and was independent of the ozonolysis reaction time. Addition of acetic acid
28 resulted in measured OH concentrations that were similar to measurements when the interference was subtracted,
29 suggesting that the source of the interference in these experiments involved the decomposition of stabilized

1 Criegee intermediates inside the IU-FAGE detection cell. Further measurements and modeling will be needed for
2 a wider variety of alkenes in order to confirm these results.

3 The interference appeared to increase with the length of the inlet in the low pressure region of the
4 detection cell, suggesting that the interference depends on the reaction time in the detection cell. However,
5 increasing the pressure in the detection cell by decreasing the flow rate did not significantly increase the observed
6 interference except for the longest inlet. This may suggest that the increase in the observed interference with the
7 length of the inlet may be the result of increased collisions of the stabilized Criegee intermediates inside the
8 detection cell leading to the formation of OH rather than the result of an increase in reaction time. Additional
9 experiments will be needed to confirm these results. To minimize this interference and other unknown
10 interferences, future measurements of OH by the IU-FAGE instrument will test different detection cell designs to
11 minimize any observed ambient interferences though changes in both reaction time and potential surface
12 collisions, while also maintaining the quality of the ambient OH measurement.

13 Regardless, future ambient measurements by the IU-FAGE instrument will incorporate the external OH
14 titration technique to quantify these and other potential unknown interferences. Because of differences in design
15 (geometry, cell pressure, flow, etc.) these interference measurements may not apply to other LIF-FAGE
16 instruments. However, it is recommended that future OH measurements using the LIF-FAGE technique
17 incorporate an external OH titration scheme or some other method to quantify potential artefacts.

18
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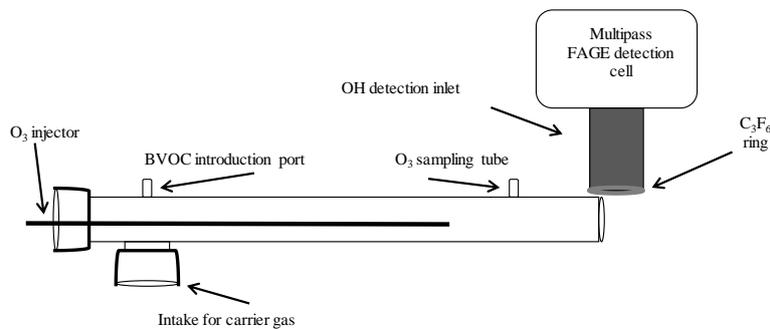


Figure 1: Schematic of the atmospheric pressure flow system used in this study.

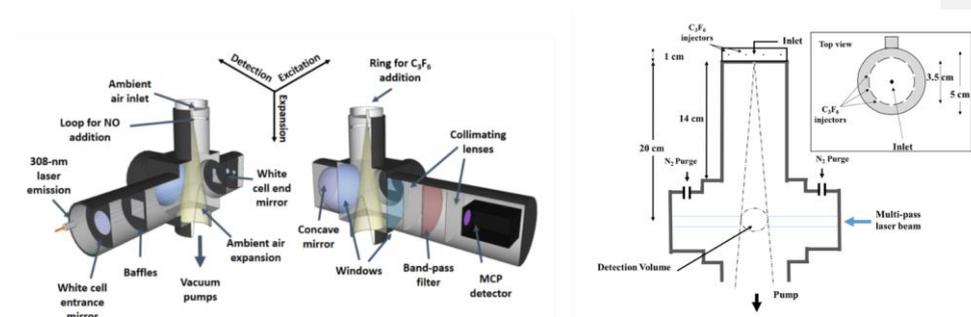


Figure 2. Schematic of the IU-FAGE sampling/excitation axis (left) and cross section with dimensions for the medium inlet configuration (right).

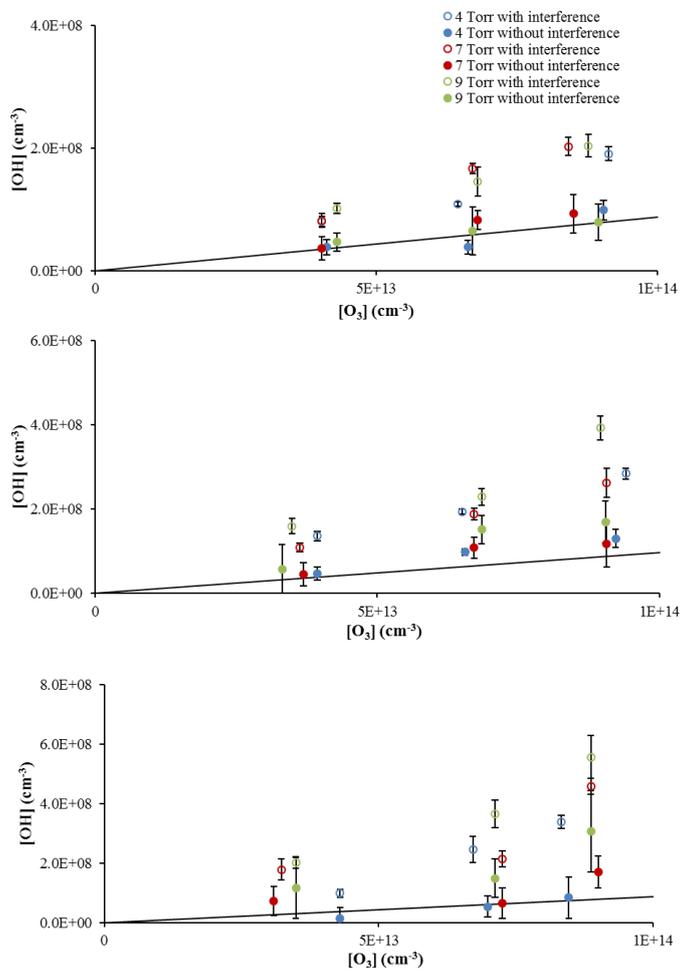


Figure 3. OH concentrations from α -pinene ozonolysis at three cell pressures with and without the interference using the 0.6 mm diameter nozzle and the short (top), medium (middle), and long (bottom) inlet lengths with α -pinene concentrations of approximately $3 \times 10^{12} \text{ cm}^{-3}$. Error bars indicate the precision of the measurement (1σ). Lines indicate the expected steady-state OH concentration based on published values of the OH yield (see text).

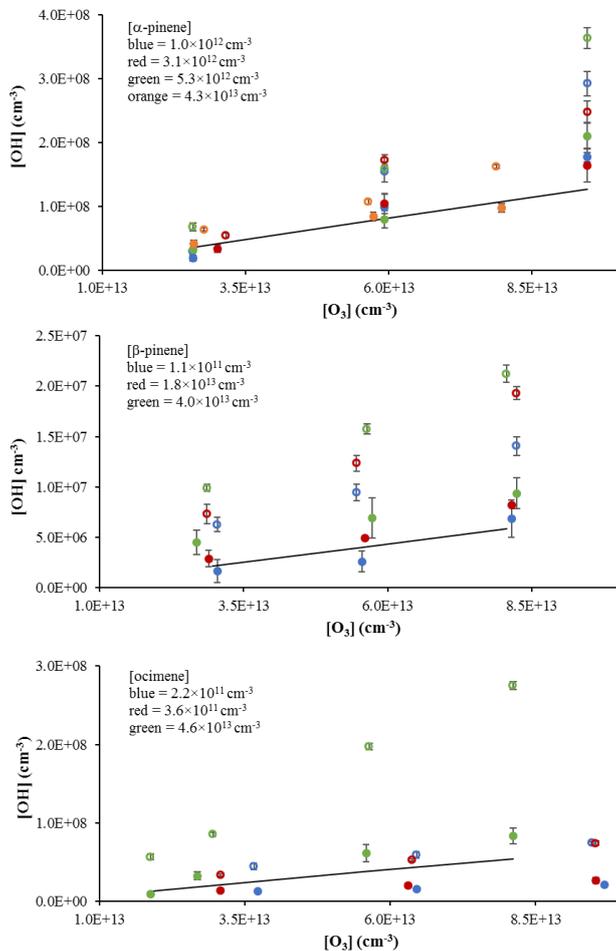


Figure 4. OH concentrations using the short inlet from the ozonolysis of α -pinene (top), β -pinene (middle), and ocimene (bottom). Open circles indicate measurements with the interference, filled circles indicate measurements without the interference. Colors indicate estimated concentrations. Solid lines reflect the expected OH radical yield (see text). Error bars reflect the measurement precision (1σ).

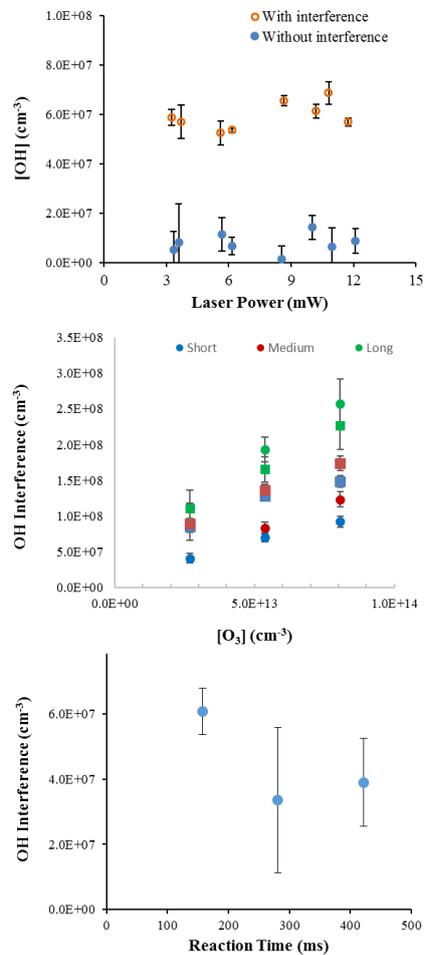


Figure 5. OH concentrations from ocimene ozonolysis with varied laser power at an ozone mixing ratio of 2 ppm and ocimene concentration of approximately $3 \times 10^{13} \text{ cm}^{-3}$ (top). OH interference during α -pinene (approximately $3 \times 10^{12} \text{ cm}^{-3}$, circles) and ocimene (approximately $3 \times 10^{13} \text{ cm}^{-3}$, squares) ozonolysis based on ozone concentration and inlet length (middle). OH interference measurements during ocimene ozonolysis as a function of reaction time with an ocimene concentration of approximately $3 \times 10^{13} \text{ cm}^{-3}$ (bottom).

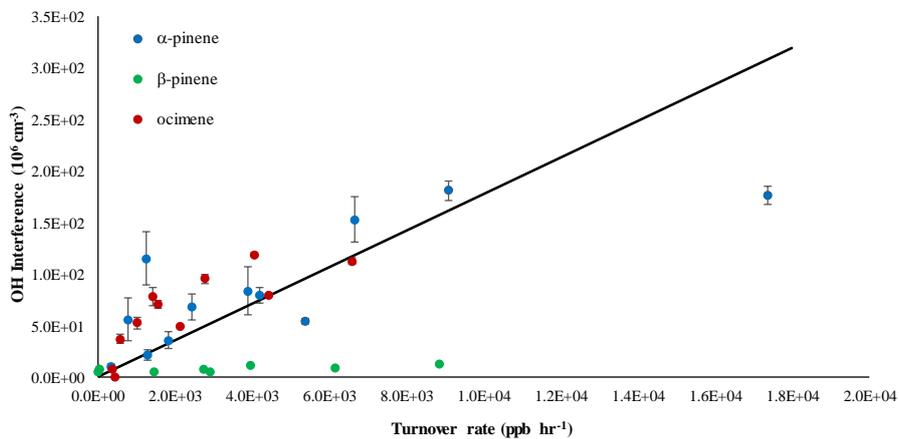


Figure 6. Interference signal as a function of turnover rate for the ozonolysis of α -pinene, β -pinene, and ocimene using the short inlet. Solid line reflects the slope observed by Fuchs et al. (2016).

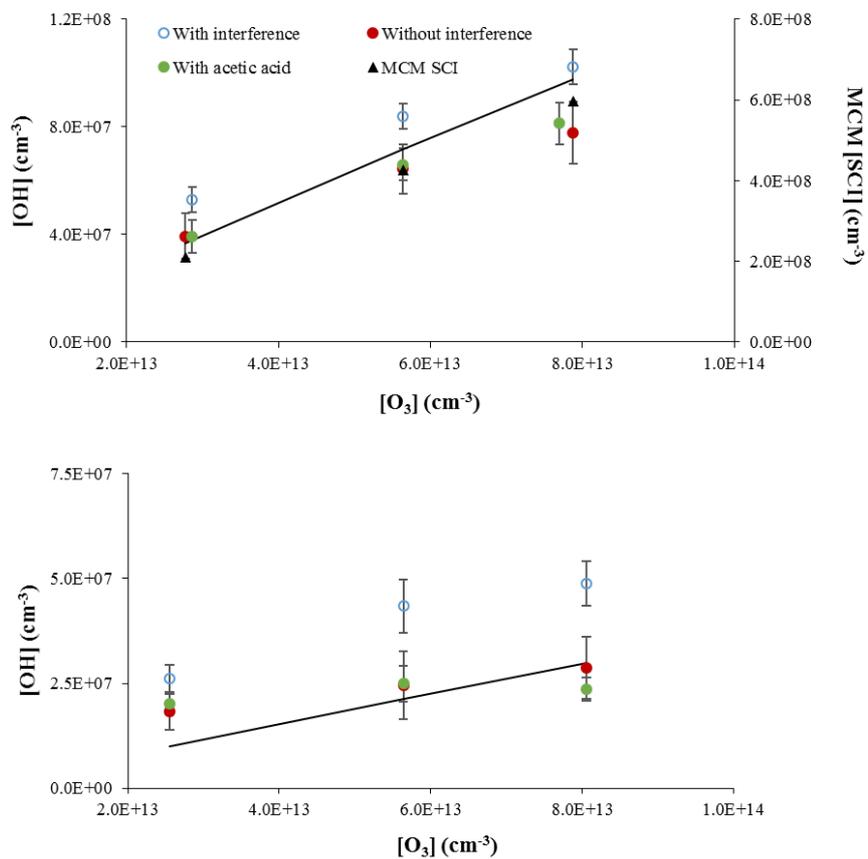


Figure 7. Measurements of OH concentrations from the ozonolysis of α -pinene (approximately $1 \times 10^{12} \text{ cm}^{-3}$, top) and ocimene (approximately $4 \times 10^{11} \text{ cm}^{-3}$, bottom). Open symbols are measurements including the interference, filled red circles are the resulting OH measurements when the interference determined by C_3F_6 addition is subtracted, and the filled green circles are the OH measurements when acetic acid is added to the flow tube. The lines indicate the expected OH concentration from published OH yields. The black triangles in the top plot reflect the predicted concentration of stabilized Criegee intermediates by the Master Chemical Mechanism (MCM SCI) (see text).